**YEAR 12 CHEMISTRY - ATCHE**

**TASK 8 PART A: PRACTICAL ASSESSMENT 2023**

**Identifying Unknown Organic Compounds**

**Part A: /12**

**Part B: /20**

**Total: /32**

**Suggested times:**

* *Part A (experimentation): 40 minutes*
* *Part B (validation): 30 minutes*

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Teacher: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Introduction:**

The identification and characterization of the structures of unknown substances are an important part of organic chemistry.

In this experiment you will carry out several **qualitative** tests that will allow you to identify functional groups in unknown organic substances. You will apply what you have learned by characterizing the unknown organic compounds in terms of their chemical reactions.

You will be given SMALL liquid samples of

**cyclohexene, ethanol, ethanoic acid and cyclohexane**

each randomly labeled as A, B, C, and D. You will also be provided with common laboratory reagents including distilled water, bromine water (Br2 (aq)), marble chips (CaCO3), potassium permanganate (KMnO4) solution and concentrated sulfuric acid.

**Important notes:**

1. Use very small amounts of the lab reagents (e.g. 1 mL) in any chemical reaction to enable them to be the limiting reactant.
2. Reagents available are: **distilled water, bromine water (Br2 (aq)), marble chips (CaCO3), potassium permanganate solution (KMnO4) and concentrated sulfuric acid.**
3. Plan to do the ‘simple’ tests to the maximum number of unknowns, leaving the more complex tests to fewer numbers of unknowns to save time.
4. You cannot use acid-base indicators or ester formation as one of your tests.
5. **If you work out the last chemical by elimination, conduct a chemical test to ensure its identity is correct**.

IN YOUR RESULTS FOLLOWING:

State all observations.

Write suitable chemical equations where possible to support your observations

Identify and state the chemical name of the unknowns.

**RESULTS (12 marks)**

**Unknown Sample A**

|  |  |
| --- | --- |
| **1st Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

If required

|  |  |
| --- | --- |
| **2nd Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

|  |  |
| --- | --- |
| **3rd Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

|  |  |
| --- | --- |
| **4th Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |
| Conclusion:  Unknown sample A is | | |

**Unknown Sample B**

|  |  |
| --- | --- |
| **1st Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

If required

|  |  |
| --- | --- |
| **2nd Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

|  |  |
| --- | --- |
| **3rd Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

|  |  |
| --- | --- |
| **4th Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |
| Conclusion:  Unknown sample B is | | |

**Unknown Sample C**

|  |  |
| --- | --- |
| **1st Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

If required

|  |  |
| --- | --- |
| **2nd Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

|  |  |
| --- | --- |
| **3rd Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

|  |  |
| --- | --- |
| **4th Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |
| Conclusion:  Unknown sample C is | | |

**Unknown Sample D**

|  |  |
| --- | --- |
| **1st Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

If required

|  |  |
| --- | --- |
| **2nd Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

|  |  |
| --- | --- |
| **3rd Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |

|  |  |
| --- | --- |
| **4th Test** |  |
| **Observation** (write **no visible reaction** if no change occurs) |  |
| **Inference** (include a chemical equation where possible) |  |
| Conclusion:  Unknown sample D is | | |

**Task 8 Part B: Validation of Practical Results**

1. Write an appropriate aim for the experiment that you conducted. (1 mark)

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1. Construct a flow chart or dichotomous key that would enable the identification of the four unknown substances supplied in part A. An example has been provided below. (2 marks)

BaSO4, NaOH, Sugar, NH4Cl

NaOH, Sugar, NH4Cl

Sugar

NaOH

BaSO4

NH4Cl

Dissolve in water

Soluble

Insoluble

Litmus paper/universal indicator

Acidic

Neutral

Basic

1. With reference to your flow chart in question 2.
   1. Briefly describe all tests that would be performed and how these would identify each substance. (4 marks)

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* 1. List two safety hazards that may arise during this experiment, for each hazard, provide a method to mitigate or minimize it. (4 marks)

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1. Two of the unknown substances provided were soluble in water. Identify these species and use your understanding of intermolecular forces to explain why they are soluble. (4 marks)

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1. Two of the unknown substances provided to you can produce a fruity smelling compound when reacted together.
   1. Identify the two substances and write a balanced chemical equation for the reaction. Ensure that you include structural formulas in your response. (4 marks)
   2. Using IUPAC nomenclature, name the main product that forms in this reaction. (1 mark)